

(UNIT-1)

03418729745(WhatsApp Groups)

SAEED MDCAT TEAM

✓ INTRODUCTION TO FUNDAMENTAL CONCEPTS OF CHEMISTRY

- PAGE 1 OF 5**



- Q.21** Number of moles of CO_2 which contains 24g oxygen is
a. 1.0 mole
b. 0.25mole
c. 0.50 mole
d. 0.75 mole
- Q.22** Complete oxidation (combustion) of one mole of an organic compound requires three mole of oxygen gas. The formula of the organic compound will be
a. CH_3CH_3
b. CH_3OH
c. $\text{CH}_3\text{CH}_2\text{OH}$
d. CH_3COOH
- Q.23** A balanced chemical equation tells us about the
a. Conditions of the reaction
b. Quantitative relationship between reactants and products
c. Rate of the reaction
d. Odour of the reactants and products
- Q.24** A limiting reactant is the one which
a. Is taken in lesser quantity in volume as compared to the other reactants
b. Gives the maximum amount of the product which is required
c. Is taken in same quantity in volume as other reactants
d. Gives the minimum amount of the product
- Q.25** Which is true about both acetic acid and oxalic acid
a. Equal molar mass
b. Same molecular formula
c. Different empirical formula
d. Equal percentage of carbon
- Q.26** Which statement is incorrect
a. One gram atom of sodium is equal to one gram of sodium
b. One mole of ion is equal to ninety-six gram of sulphate
c. One gram molecule of water is equal to eighteen gram of water
d. One formula unit mass of lime stone is equal to hundred gram of lime stone
- Q.27** CH_2O is the empirical formula of
a. Formic acid
b. Sucrose
c. Lactic acid
d. Ethanol
- Q.28** Which of the following is not concentration unit of solution
a. %age composition
b. Mole fraction
c. Molarity
d. Molar volume
- Q.29** Combustion analysis is used to determine the _____ of an organic compound in i.e. hydrocarbons
a. Molecular formula
b. Empirical formula
c. Structural formula
d. Lewis structure
- Q.30** 20g NaOH is dissolved per dm^3 of the aqueous solution, its molarity will be
a. 0.5M
b. 1.0M
c. 0.25M
d. 0.1M
- Q.31** Which of the following terms is used for 238g of uranium
a. 1g molecule
b. 1g atom
c. 1g ion
d. 1g formula
- Q.32** 23g of sodium and 24g of magnesium have equal _____ in them
a. Mass
b. Number of protons
c. Number of atoms
d. All of these
- Q.33** Total number of electrons present in 34 g OH^- are



- a. $9 N_A$
c. $10 N_A$
- Q.34 Molarity of 25% (w/v) NaOH solution is**
a. 2.5 M
c. 5.0 M
- Q.35 720g of glucose contains how many moles of glucose**
a. 2
c. 4
- Q.36 Number of moles of solute dissolved per dm^3 of solution is called**
a. Molarity
c. Molality
- Q.37 The number of atoms in one gram atom of an element is**
a. N_A of atoms
c. N_A of molecules
- Q.38 An unknown compound has empirical formula CH_3O . Its molar mass is 62g/mole. The compound may be**
a. $\text{CH}_2(\text{OH})\text{CH}(\text{OH})\text{CH}_2(\text{OH})$
c. $\text{CH}_2(\text{OH})\text{CH}_2(\text{OH})$
- Q.39 For stoichiometry calculations, we have to assume**
a. Mass of reactants is less than the mass of products
c. All the reactants are completely converted into products
b. Side reaction occurs
d. Reaction is reversible
- Q.40 For more than two components solution, we prefer units of concentration like**
a. Mole fraction
c. Molality
- Q.41 Quantitative relationship between reactants and products in a balanced chemical equation is known as**
a. Stoichiometry
c. Titrimetry
- Q.42 What will be volume of 2.5 moles of chlorine molecules occupy at STP.**
a. 22.414dm^3
c. 50.207dm^3
- Q.43 Eight gram of methanol contain.**
a. 4 gram "C" 3g "H" and 1g "O"
c. 3 gram "C" 1g "H" and 4g "O"
- Q.44 2.5 mole of NH_3 and 2.5 mole of SO_3 at same temperature and pressure have equal.**
a. Volume
c. Atoms
- Q.45 The efficiency of a chemical reaction can be determined with the help of _____**
a. Actual yield
c. Theoretical yield
- Q.46 While finding the relative atomic mass, which of the following standard is used to compare the atomic mass of chlorine (35.5 amu)**
a. Neon-20
c. Nucleon number
- Q.47 What is the mass of Al in 204g of Al_2O_3 ?**
a. 126g
c. 108g
- Q.48 One mole of ethanol and one mole of ethane have equal**
a. Mass
c. Number of electrons
- Q.49 58.5amu mass of NaCl called as**
a. Molecular mass
c. Ionic mass
- b. $18 N_A$
d. $20 N_A$
- b. 1.5 M
d. 6.25 M
- b. 6
d. 8
- b. Mole fraction
d. ppm
- b. N_A of ions
d. N_A of formula unit
- b. CH_3COCH_3
d. $\text{C}_2\text{H}_5\text{OH}$
- b. Molarity
d. ppm
- b. Spectrometry
d. Chromatography
- b. 56.035cm^3
d. 56.035dm^3
- b. 3 gram "C" 2g "H" and 3g "O"
d. 2 gram "C" 1g "H" and 5g "O"
- b. Molecules
d. All of these
- b. Expected yield
d. %age yield
- b. Carbon-13
d. Carbon-12
- b. 27g
d. 54g
- b. Number of atoms
d. Number of molecules
- b. Formula mass
d. Atomic mass



- Q.50** The amount of product obtained without performing an experiment is known as
- a. Expected yield
 - b. Calculated yield
 - c. Theoretical yield
 - d. All of these

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